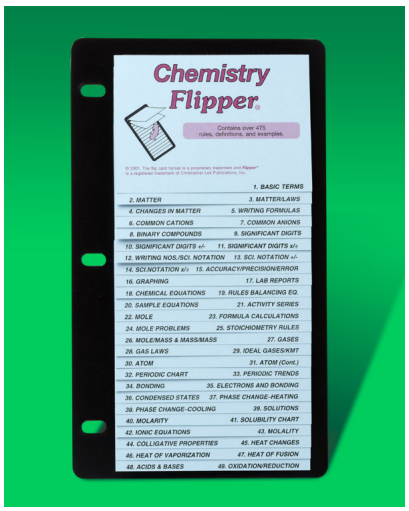
Chemistry



Chemistry Flipper®-Topics - CLP-590W Introduction to Chemistry. Grades 10–College.

Rating: Not Rated Yet
Price
Sales price $9.95

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Description

Introduction to Chemistry. Grades 10–College.

Chemistry Flipper®-Topics - CLP-590W

- Basic Terms
- Basic Terms: Observation
- Basic Terms: Qualitative
- Basic Terms: Quantitative
- Basic Terms: Hypothesis
- Basic Terms: Inference
- Basic Terms: Prediction
- Basic Terms: Variable
- Basic Terms: Control
- Basic Terms: Control Setup
- Basic Terms: Theory
- Basic Terms: Data
- Basic Terms: Model
- Basic Terms: Property
- Basic Terms: Intensive

- Formula Calculations: Determine an Empirical Formula
- Formula Calculations: Determine a Molecular Formula
- Working Mole Problems
- Working Mole Problems: Divide in, Multiply Out
- Stoichiometry Rules
- Stoichiometry Rules: How to Work Problems
- Mole/Mass Problems
- Mass/Mass Problems
- Gases
- Gases: Pressure
- Gases: Temperature
- Gases: Dalton's Law of Partial Pressures
- Gases: Avogadro's Principle
- Gas Laws
- Gas Laws: Charles's Law
- Gas Laws: Boyle's Law
- Gas Laws: Law of Combining Volumes
- Ideal Gas Law
- Assumptions of the Kinetic Molecular Theory
- Atom
- Atom: Atomic Number
- Atom: Mass Number
- Atom: Atomic Mass
- Atom: Nucleus
- Atom: Nucleus: Proton
- Atom: Nucleus: Neutron
- Atom: Electron
- Atom: Electron Configuration
- Atom: Isotopes
- Atom: Ion
- Atom: Ion: Cation
- Atom: Ion: Anion
- Periodic Chart
- Periodic Chart: Group/Family
- Periodic Chart: Period
- Periodic Chart: Metals
- Periodic Chart: Nonmetals
- Periodic Chart: Metalloids
- Periodic Chart: Halogens
- Periodic Chart: Noble Gases
- Periodic Law
- Periodic Law
- Periodic Trends: Atomic Radius
- Periodic Trends: Ionic Radii
- Periodic Trends: First Ionization Energy
- Bonding

- Basic Terms: Extensive
- Basic Terms: Heterogeneous
- Basic Terms: Homogeneous
- Basic Terms: Macroscopic
- Basic Terms: Microscopic
- Matter
- Matter: Mass
- Matter: Volume
- Matter: Density
- Matter: Derived Quantity
- Matter: Pure Substance
- Matter: Element
- Matter: Atom
- Matter: Compound
- Matter: Bonding
- Matter: Molecule
- Matter: Mixture
- Matter: Solution
- Matter: Solution: Solute
- Matter: Solution: Solvent
- Matter: Fluid
- Matter: Physical States of Matter
- Matter: Physical States of Matter: Solid
- Matter: Physical States of Matter: Liquid
- Matter: Physical States of Matter: Gas
- Matter: Physical States of Matter: Plasma
- Laws
- Laws: Conservation of Mass
- Laws: Conservation of Energy
- Laws: Definite Composition
- Laws: Multiple Proportions
- Changes in Matter
- Changes in Matter: Energy
- Changes in Matter: Energy: Potential Energy
- Changes in Matter: Energy: Kinetic Energy
- Changes in Matter: Chemical Change
- Changes in Matter: Chemical Reaction
- Changes in Matter: Physical Change
- Changes in Matter: Phase Change
- Changes in Matter: Nuclear Change
- Changes in Matter: Exothermic
- Changes in Matter: Endothermic
- Writing Formulas
- Writing Formulas: Chemical Formula
- Writing Formulas: Chemical Formula: Subscripts
- Writing Formulas: Binary Compounds
- Writing Formulas: Diatomic Elements
- Writing Formulas: Polyatomic Elements
- Common Cations
- Common Cations: Listing
- Common Anions
- Common Anions: Listing
- Naming Binary Compounds
- Rules for Significant Digits
- Rules for Significant Digits: Nonzero Numbers
- Rules for Significant Digits: Leading Zeros
- Rules for Significant Digits: Capitve Zeros
- Rules for Significant Digits: Trailing Zeros
- Rules for Significant Digits: Numbers in Scientific Notation
- Rules for Significant Digits: Zeroes With a Line Over Them
- Significant Digits +/-
- Significant Digits x/÷
- Writing Numbers in Scientific Notation
- Writing Numbers in Scientific Notation: Base
- Writing Numbers in Scientific Notation: Exponent
- Writing Numbers in Scientific Notation: Significant Digits
- Writing Numbers in Scientific Notation: Change to Decimal Form
- Scientific Notation +/-

- Bonding: Chemical Bonds
- Bonding: Ionic Bond
- Bonding: Covalent Bond
- Bonding: Nonpolar Covalent
- Bonding: Polar Covalent Bond
- Bonding: Electronegativity
- Representing Electrons and Bonding
- Representing Electrons and Bonding: Octet Rule
- Representing Electrons and Bonding: Electron Dot Symbols
- Representing Electrons and Bonding: Structural Formulas
- Condensed States of Matter
- Condensed States of Matter: Intermolecular Forces
- Condensed States of Matter: Intermolecular Forces: Dispersion Forces
- Condensed States of Matter: Intermolecular Forces: Dispersion Forces: Molecules
- Condensed States of Matter: Intermolecular Forces: Dipole-Dipole Forces
- Condensed States of Matter: Intermolecular Forces: Hydrogen Bonding
- Condensed States of Matter: Liquids
- Condensed States of Matter: Liquids: Surface Tension
- Condensed States of Matter: Liquids: Water
- Condensed States of Matter: Solids
- Phase Change - Heating
- Phase Change - Cooling
- Solutions
- Solutions: Solute
- Solutions: Solvent
- Solutions: Solvation
- Solutions: Hydration
- Solutions: Dissociation
- Solutions: Precipitation
- Solutions: Miscibility
- Solutions: Solubility
- Solutions: Solubility: Saturated
- Solutions: Solubility: Unsaturated
- Solutions: Solubility: Supersaturated
- Solutions: Factors Affecting Solubility
- Molarity
- Solubility Chart
- Solubility Chart: Soluble
- Solubility Chart: Insoluble
- Ionic Equations
- Ionic Equations: Spectator Ions
- Ionic Equations: Net Ionic Equation
- Ionic Equations: Write an Ionic Equation
- Molality
- Colligative Properties
- Colligative Properties: Boiling Point Elevation
- Colligative Properties: Freezing Point Depression
- Heat Changes
- Heat Changes: Molar Heat
- Heat Changes: Molar Heat: Exothermic Reaction
- Heat Changes: Molar Heat: Endothermic Reaction
- Heat Changes: Specific Heat
- Heat of Vaporization
- Heat of Fusion
- Acids and Bases
- Acids and Bases: Arrhenius Definitions
- Acids and Bases: Bronsted Definitions
- Acids and Bases: Properties
- Acids and Bases: Properties: Taste
- Acids and Bases: Properties: Feel
- Acids and Bases: Properties: Reaction With Metals
- Acids and Bases: Properties: Electrical Conductivity
- Acids and Bases: Properties: Indicators
- Acids and Bases: Properties: Indicators: Litmus
- Acids and Bases: Properties: Indicators: Phenolphthalein
- Acids and Bases: Conjugate Acid-Base Pair
- Acids and Bases: Neutralization
- Acids and Bases: pH
- Oxidation/Reduction
- Oxidation/Reduction: Oxidizing Agents
- Oxidation/Reduction: Reducing Agents
- Oxidation/Reduction: "Redox" Reactions
- Oxidation/Reduction: Oxidation Number
- Oxidation/Reduction: Rules for Assigning Oxidation Numbers

- Scientific Notation +/-: Significant Digits
- Scientific Notation x/÷
- Scientific Notation x/÷: Significant Digits
- Accuracy
- Absolute Error
- Percent Error
- Precision
- Deviation
- Exact Number
- Graphing
- Graphing: Direct Relationship
- Graphing: Indirect (Inverse) Relationship
- Lab Reports
- Lab Reports: Purpose
- Lab Reports: Materials
- Lab Reports: Method/Procedure
- Lab Reports: Observations
- Lab Reports: Calculations
- Lab Reports: Results/Discussion
- Lab Reports: Conclusion
- Chemical Equations
- Chemical Equations: Listing
- Rules for Balancing Equations
- What a Balanced Equation Shows
- Sample Equations
- Sample Equations: Hydrogen and Iodide
- Sample Equations: Hydrogen Bromide and Oxygen
- Activity Series of the Elements
- Activity Series of the Elements: Listing of Metals
- Activity Series of the Elements: Listing of Halogens
- Mole
- Mole: Avogadro's Number
- Mole: Mole of Particles
- Mole: Molar Mass
- Formula Calculations
- Formula Calculations: Determine the Percent Composition of a Compound